

This is the accepted manuscript made available via CHORUS. The article has been published as:

## High Pressure Stabilization and Emergent Forms of $\text{PbH}_{\{4\}}$

Patryk Zaleski-Ejgierd, Roald Hoffmann, and N. W. Ashcroft

Phys. Rev. Lett. **107**, 037002 — Published 11 July 2011

DOI: [10.1103/PhysRevLett.107.037002](https://doi.org/10.1103/PhysRevLett.107.037002)

# The very heavy and the very light: high pressure stabilization and emergent forms of $\text{PbH}_4$

Patryk Zaleski-Ejgierd<sup>1</sup>, Roald Hoffmann<sup>2</sup> and N. W. Ashcroft<sup>1</sup>

<sup>1</sup> *Laboratory of Atomic and Solid State Physics and Cornell Center for Materials Research, Clark Hall, Cornell University, Ithaca, New York 14853-2501, USA*

<sup>2</sup> *Department of Chemistry and Chemical Biology, Baker Laboratory, Cornell University, Ithaca, New York 14853-1301, USA*

## Abstract

A wide decomposition pressure range of 132 GPa is predicted for  $\text{PbH}_4$  above which it emerges in very different forms compared with its Group 14 congeners. This triply Born-Oppenheimer system is a non-molecular, three-dimensional, metallic alloy, despite a prominent layered structure. A significant number of enthalpically near degenerate structures, with exceedingly small energy barriers for distortions, and characteristic instabilities in the phonon spectra suggest that even at very high pressures  $\text{PbH}_4$  may exhibit both metallic and liquid-like properties, and sub-lattice or even full melting.

Among Group 14 hydrides, the lead-hydrogen system is special, for it clearly combines one of the heaviest (Pb) elements with the lightest (H). The attendant mass ratio (207:1) immediately implies a distinct separation of time-scales for the ensuing dynamics, in

very much the way that standard Born-Oppenheimer separation commonly applies to the electrons and both nuclei (Pb and H).

It has been suggested that Group 14 hydrides, including methane ( $\text{CH}_4$ ), silane ( $\text{SiH}_4$ ), germane ( $\text{GeH}_4$ ) and stannane ( $\text{SnH}_4$ ) might become metallic under pressures lower than is projected for pure  $\text{H}_2$  metallization [1]. Experimental attempts to metallize methane, the “lightest” of the hydrides, have so far failed. Silane passes through phases of increasing coordination, but despite initial findings of metallization, appears to remain non-metallic to 150 GPa [2, and references therein]. For germane [3, 24] and stannane [17] metallization is predicted at a variety of accessible pressures; in both systems the high pressure phases interestingly contain paired  $\text{H}_2$  units.

Chemically,  $\text{PbH}_4$  still remains the most elusive of Group-14 tetrahydrides. The first attempts to synthesize  $\text{PbH}_4$ , employing the classical methods used to synthesize  $\text{GeH}_4$  or  $\text{SnH}_4$ , failed [4, 5, 6, 7]. The pioneering theoretical work of Pyykkö *et al.* utilized Dirac-Hartree-Fock relativistic calculations to predict the structure and stability of  $\text{PbH}_4$  [8, 9]. A predicted tetrahedral structure of an isolated molecule, with an equilibrium Pb-H distance of approx. 1.73 Å, was eventually confirmed by experiment [10, 11]. Krivtsun *et al.* noted that the plumbane molecules are kinetically unstable and readily decompose to a metallic lead film and  $\text{H}_2$  in approximately 10 s [7]. The high reactivity and short life-time of  $\text{PbH}_4$  have so far combined to prevent any crystallographic investigation of the species under normal conditions. We have therefore undertaken here a theoretical analysis of its structural, thermodynamic, and electronic properties, particularly at elevated densities.

As starting points and guides in a structure search we used a number of geometries previously reported for Group-14 tetrahydrides,  $\text{XH}_4$ , as well as other structures in

which lead is 2- to 16- fold coordinated by hydrogen, and with up to  $Z = 4$  formula units per cell. We scanned 1000 randomly generated structures at the selected pressures and these explorations were complemented by the use of the USPEX evolutionary algorithm [12, 13] to seek the most stable of these. Finally, for the lowest enthalpy structures we also investigated the nuclear dynamics through calculated phonon spectra using the PHON code [14] later pursuing any ensuing imaginary modes to more stable and secure minima. The methodology is detailed in the Supplementary Material (SM) for this paper.

Nine energetically competitive structures arise from this search (see Fig. S1 of the SM). They all show a surprisingly convergent evolution of the enthalpy, by which we mean that a number of quite different ground state structural possibilities, each corresponding to a meta-stable minimum, are enthalpically close to each other over a wide range of pressures (see Fig. 1). Below the calculated decomposition pressure (132 GPa),  $\text{PbH}_4$  forms extended layered structures rather than crystalline forms with no vestige of distinct plumbane molecules. Were we to proceed to even larger unit cells (we are currently constrained to  $Z=2$  and  $Z=4$ ) we are confident that for  $P < 132$  GPa the layered structures we will find would segregate further towards the  $\text{Pb} + 2\text{H}_2$  limit. The computed  $\Delta H$  of a single gas-phase molecule is 2.74 eV, so this large decomposition pressure range comes as no surprise.

For  $P > 132$  GPa at least seven distinct phases prove to be more stable than the separated elements. Per formula unit all of them remain within approx. 0.1-0.2 eV of each other over the large pressure range, with the exception of Phase I. The two most stable structures, Phase VII (Imma) and VIII (Ibam), are representative, and singled out

for more detailed discussion. Figure 2 shows their structure at 300 GPa, and Table I provides some salient structural parameters.

In Phase VII, atoms of lead contribute to the binding of two neighboring layers with relatively short bonds of 2.64 Å at 300 GPa, these closely similar to the Pb-Pb bond distances within the lead layers (see Table 1). The 3D lead framework has 1D-like infinite channels filled by H<sub>2</sub> units these coupled in such way that they form flat (H<sub>2</sub>)<sub>2</sub> quartets. At even higher pressures, increasing coordination is developing; at 300 GPa the shortest calculated intermolecular H<sub>2</sub>-H<sub>2</sub> separation, or contact, in the (H<sub>2</sub>)<sub>2</sub> quartet is 1.07 Å, while the shortest distance between the parallel quartets is 1.36 Å.

**Table I. Selected bond lengths (Å), number of Pb-H contacts\* and  $r_s$  values calculated for selected phases of PbH<sub>4</sub> at 300 GPa.**

| Phase | Pb-Pb       |               | H-H  | H <sub>2</sub> -H <sub>2</sub> <sup>a</sup> | Pb-H   | N of contacts<br>Pb-H < 2.30 Å | $r_s$ |
|-------|-------------|---------------|------|---|--|--------------------------------|-------|
|       | inter-layer | intra-layer   |      |   |  |                                |       |
| VII   | 2.64        | 2.55/<br>2.67 | 0.79 | 1.07/1.36                                   | 1.90(x2), 1.95(x4), 2.10(x2)<br>2.12(x4), 2.15(x4) | 16                             | 1.56  |
| VIII  | 3.06        | 2.62/<br>2.71 | 0.81 | 1.02/1.16                                   | 1.88(x4), 2.03(x4), 2.17(x4)<br>2.18(x4)           | 16                             | 1.55  |

\* selected shortest Pb-H distances, <sup>a</sup> shortest H-H separation

At the highest analyzed pressures Phase VIII is most stable with quasi-hexagonal one-atom thick layers of lead intercalated by layers of hydrogen. All hydrogen atoms are co-planar, even though *not* constrained to be so during structural optimization. Within these planar hydrogen layers, the (H<sub>2</sub>)<sub>2</sub> quartets remain distinctively present, as Fig. 2 shows. Once again, higher dimensionality for the hydrogen framework looms: at 300 GPa the shortest H<sub>2</sub>-H<sub>2</sub> contact is 1.02 Å, while the next shortest contact is only

1.16 Å. The  $(\text{H}_2)_2$  quartets can be viewed in such way that they actually form 10-membered rings, as illustrated in Fig. 2. A somewhat similar layered structure has been observed by Martinez-Canales *et al.* for high-pressure silane (Pbcn,  $P > 220$  GPa). In it, slightly displaced square Si layers were intercalated by two-atom thick layers of hydrogen.[15] In contrast to  $\text{PbH}_4$ , the intra- and inter-molecular H-H distances are far less equalized, at 0.74 and 1.35 Å, respectively.

How shall we look at the unusual bonding in the  $\text{H}_2$  layers in these  $\text{PbH}_4$  structures? First, the initially (low P) unbound hydrogens appear to be closer to each other than in  $\text{H}_2$  itself. At 300 GPa the shortest *inter*-molecular distances (Table 1) are 1.07 Å (Phase VII), and 1.02 Å (Phase VIII). In the most stable  $\text{H}_2$  structure at this pressure (Cmca-12), the corresponding distance is 1.10 Å [16]. Second, the *intra*-molecular  $\text{H}_2$  distance is notably longer in the  $\text{PbH}_4$  structures – 0.79 Å (Phase VII), and 0.81 Å (Phase VIII) vs. 0.75-0.78 Å in  $\text{H}_2$  (Cmca-12 at 300 GPa).

There may be two reasons behind this behavior: a) the *effective* pressure on the hydrogen sub-lattice might be greater in  $\text{PbH}_4$  than in  $\text{H}_2$ . Here the Pb's could well be thought of as now present in less compressible layers that reduce the effective volume (and concomitantly increase the effective pressure) on the  $\text{H}_2$ 's. b) there may be some 'squeezing out' of electron density from the Pb layers at high pressures, impelling electrons into effective  $\text{H}_2$   $\sigma_u^*$  levels [see also ref. 17].

A second important aspect of the hydrogen sub-lattice, one already noted, is that all of our proposed structures contain hydrogen units with a strongly expressed tendency for  $\text{H}_2$ - $\text{H}_2$  pairing. That this is not just an imagined association may be seen in Fig. 3, where we show the Electron Localization Function (ELF) [18] calculated for the Phase VIII at 400 GPa. The small value of the ELF between two neighboring atoms in the

direction perpendicular to lead layers supports the interpretation of the layered character of Phase VIII. The enhanced value of ELF between the H atoms of a single H<sub>2</sub> unit is expected, but note that the ELF also indicates significant bonding between *two* H<sub>2</sub> units, forming H<sub>2</sub> quartets, and a decrease of ELF outside of the quartets.

Because of the high atomic number of lead, relativistic effects must be taken into account. Scalar relativistic (SR) and spin-orbit (SO) effects are both of importance. Christensen showed that non-relativistic lead would adopt the diamond-type structure rather than fcc [19], while scalar-relativistic lead would preserve the fcc lattice. In our calculation the SR effects are incorporated through the use of pseudo-potentials (PP); they account for the major part of the influence of relativity. We have included spin-orbit effects in a series of trial calculations, but have found them to have small impact on the enthalpy *differences* though they are well known to influence the Fermi surface of lead (and also in white tin).

Starting from 25 GPa all phases behave almost identically upon compression, hence the equation of state (EOS), which connects different phases as the pressure increases, is notably smooth (for the EOS computed at T=0 and for a static lattice see Fig. S4).

Over a wide range of pressures a number of more-or-less layered structures, some structurally quite different, are actually quite close in enthalpy (Fig. 1). The zero point energy of each of the structures (not included in the curves displayed) is quite comparable to the difference in enthalpy of the various structures. This situation is suggestive of eventual hydrogen sub-lattice melting at elevated pressures or even of diffusive or liquid-like ground states. In fact, perhaps a way to think about these structures is as layers of lead effectively 'lubricated' by highly mobile intercalated layers of hydrogen molecules. The latter in turn begin to coalesce into quartets and an

extended layer under pressure. How might we therefore probe a tendency towards liquid or partial liquid behavior, possibly even a liquid crystal in  $\text{PbH}_4$ ? One direction is to examine the dynamics as revealed in phonon modes of these structures, in particular as part of a search for low frequency transverse modes indicative of easy shearing.

Phonon calculations for the hypothetical  $\text{PbH}_4$  phases are of importance especially as guarantors of their stability but also with respect to the information such calculations may yield on easy motions in the structures. They are also of considerable interest for assessment of possible superconductivity by the phonon mechanism in metallic modifications of  $\text{PbH}_4$  (lead already being known as a strong coupling superconductor, and the high hydrogen modes plausibly assisting further, as mooted for pure metallic hydrogen (and deuterium)). Within the pressure range of Fig. 1, all the structures exhibit only real frequencies at the  $\Gamma$ -point. All show clearly separated hydrogen vibron modes located in the band  $3500\text{-}4000\text{cm}^{-1}$  for phases where the  $\text{H}_2$  molecules are relatively unperturbed; they are  $\sim 500\text{cm}^{-1}$  lower when the  $\text{H}_2$  molecules are elongated (Phases VII and VIII). Nicely distinguishable low frequency modes are also associated with lead framework displacements.

Phase VIII is clearly interesting because of its pronounced layered character. In fact, we find it to be the most stable phase at pressures higher than 300 GPa. The full phonon dispersion spectrum and the corresponding phonon density of states (PhDOS) per atom (and per  $\text{cm}^{-1}$ ) calculated at 400 GPa is shown in the SM. Upon increase of the displacements (used to calculate the phonon modes) our calculations eventually yield imaginary frequencies. Attempts to follow these necessarily destabilizing modes result only in a restoration of the original structure. Analysis of the modes indicates that they correspond to transverse motions within the hydrogen sub-lattices, accompanied by



negligible buckling of lead layers. Increase of pressure makes the imaginary modes less prominent. Deuterium also stabilizes the imaginary modes; using tritium enhances this mass effect further.

The fact that the layered Phase VIII cannot support high-amplitude transverse phonons suggests that it has a very low shear modulus, or that the hydrogen sub-lattices may even exhibit a 2D liquid-like character (sub-lattice melting or diffusion). Given this, we return to the point concerning the very significant differences in time scales between the lead and hydrogen subsystems. In the Born-Oppenheimer approximation this suggests that for temporarily fixed lead coordinates, the hydrogenic states may be integrated- or traced-out in favor of effective but purely Pb-Pb types of interactions. Since the hydrogens are appearing between Pb-Pb layers, this in turn suggests that (after averaging over the local phase space of hydrogen positions) the net interactions may be notably reduced (a form of ‘screening’ is taking place).

In this context we examine one of the standard measures (the ‘small parameter’) associated with the adiabatic separation of time-scales. For an electron-proton system it is  $(m_e/m_H)^{(1/4)}$  or  $\sim 0.15$  [20]. For the elements the Born-Oppenheimer approximation is least well satisfied for hydrogen, but is still generally adequate. For the hydrogen-lead system it would be  $(m_H/m_{Pb})^{(1/4)}$  which is just  $(1/207)^{(1/4)} \sim 0.26$  yielding an interesting comparison with the electron-hydrogen case. Finally, for the electron-lead system it will be  $(m_e/m_{Pb})^{(1/4)} \sim 0.04$  with the corresponding time scales then being exceedingly well separated, as is expected.

It is of some considerable interest that all of the reported structures are computed to be metallic, and strongly so, as determined by the band-structures and densities-of-states. The metallic character is found over the whole studied pressure range and in fact

at high pressures  $\text{PbH}_4$  exhibits properties of a nearly free electron metal. By way of example we present in Figure 4 the electronic density of states (DOS) calculated for Phases VII and VIII at 300 GPa (the corresponding band structures are in the SM). We emphasize that the overall DOS is very much free-electron like in a three dimensional context; in the DOS there is no sign of the gap (as in the case of  $\text{CH}_4$ ) or pseudo-gap found in calculations on other  $\text{EH}_4$  systems ( $\text{E} = \text{Si}, \text{Ge}, \text{Sn}$ ) [14, 21, 22, 23, 24, 25, 26].

Detailed analysis shows that hydrogen states contribute to the DOS throughout this band (see SM). These observations indicate, therefore, that at high-pressure  $\text{PbH}_4$  might reasonably be viewed as an alloy of metallic lead and metallic hydrogen. Note also that there is no sign in the total DOS of any rectangular onsets, the hallmark of two-dimensionality at low energy. While for phase VIII the structure appears geometrically layered, it is *decidedly* three-dimensional in electronic terms.

In any attempt to fabricate  $\text{PbH}_4$  under pressure, starting for instance with ultra-divided lead and liquid hydrogen, it is important to be aware of the presence of other possibly competing stoichiometries, for example  $\text{PbH}_2$ ,  $\text{PbH}_6$  and  $\text{PbH}_8$ . We will report on an exploratory search for structures of these separately. But briefly, each also has a region of stability, and at high pressures  $\text{PbH}_x$  ( $x = 6, 8$ ) are quite stable and also metallic.

For  $\text{PbH}_4$ , the heaviest Group 14 tetra-hydride, decomposition into the elements over a wide pressure range ( $0 < P < 132$  GPa) is indeed first predicted. A solid containing the molecular plumbane form,  $\text{PbH}_4$ , analogous to  $\text{CH}_4$ , is *never* stable thermodynamically. But above the decomposition pressure we find a large number of

energetically almost indistinguishable local minima. One of these is Phase VII (Imma) and in it a three-dimensional lead framework is packed with one-dimensional hydrogen chains composed of characteristic  $(\text{H}_2)_2$  dimers. Above 296 GPa Phase VIII (Ibam) is found to be the most stable; in it layers of lead are intercalated by parallel layers of hydrogen but the ensuing electronic structure is notably three dimensional.

From the standpoint of possible experiments it also may be important to note that both phases are metallic at *any* of the studied pressures and the density of states at the Fermi level is close to the free electron value. And the H-H distances in these structures are more equalized than those in pure  $\text{H}_2$ . It appears that the effective pressure for the hydrogens is greater than the nominal one but despite the layer-like geometry of phase VIII (the most stable) we have for this hypothetical phase a three-dimensionally delocalized alloy-like electronic structure.

Thus the present study shows that external pressure can notably stabilize  $\text{PbH}_4$ , one of the heaviest elements now combined with the lightest. A significant number of enthalpically near degenerate structures, with small energy barriers for distortions and characteristic instabilities in the phonon spectra also suggest that even at very high pressures  $\text{PbH}_4$  may exhibit not only metallic but also diffusive or liquid-like properties.

We thank the National Science Foundation (DMR-0907425) for financial support. Part of this research was carried out in conjunction with EFree, an Energy Frontier Research Center funded by the U.S. Department of Energy, Office of Science, Office of Basic Energy Sciences under Award Number DESC0001057.

**Fig. 1.** (Color online) Calculated enthalpy difference curves for  $\text{PbH}_4$ . The combined enthalpies of the elements in their most stable form [27, 28] establish the reference (zero) line. Note both the smoothness and also the significant overlap of the curves over a wide range of pressures.

**Fig. 2.** (Color online). The two most stable high-pressure structures of  $\text{PbH}_4$ . (Left) Phase VII at 250 GPa; the 3D lead framework is noteworthy. (Right) Phase VIII at 400 GPa; it is strikingly layered.

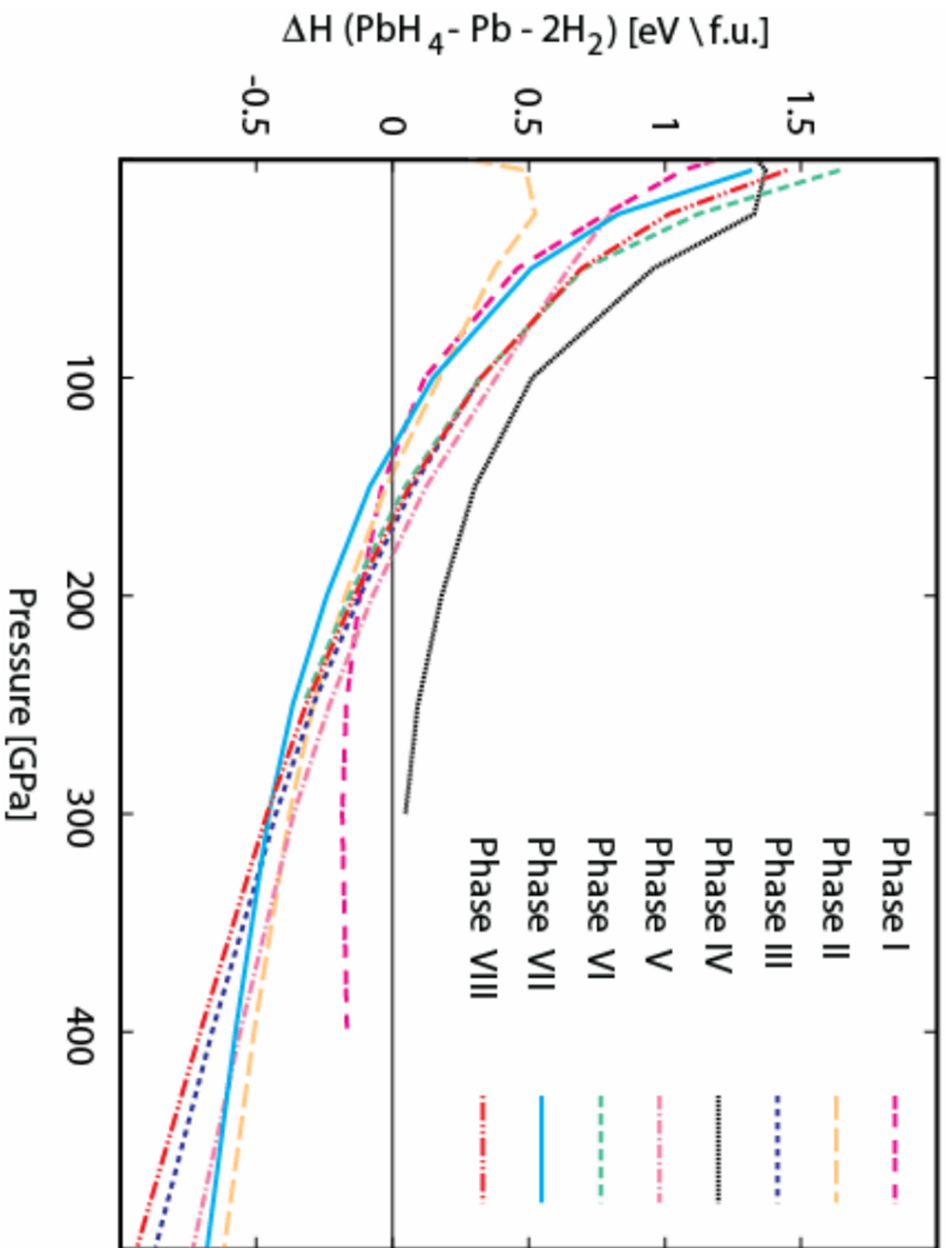
**Fig. 3.** The Phase VIII structure of  $\text{PbH}_4$  with three superimposed electron localization function (ELF) cross-sections, calculated at 400 GPa. Note the high value of the ELF between four hydrogen atoms in  $(\text{H}_2)_2$  quartets (see text).

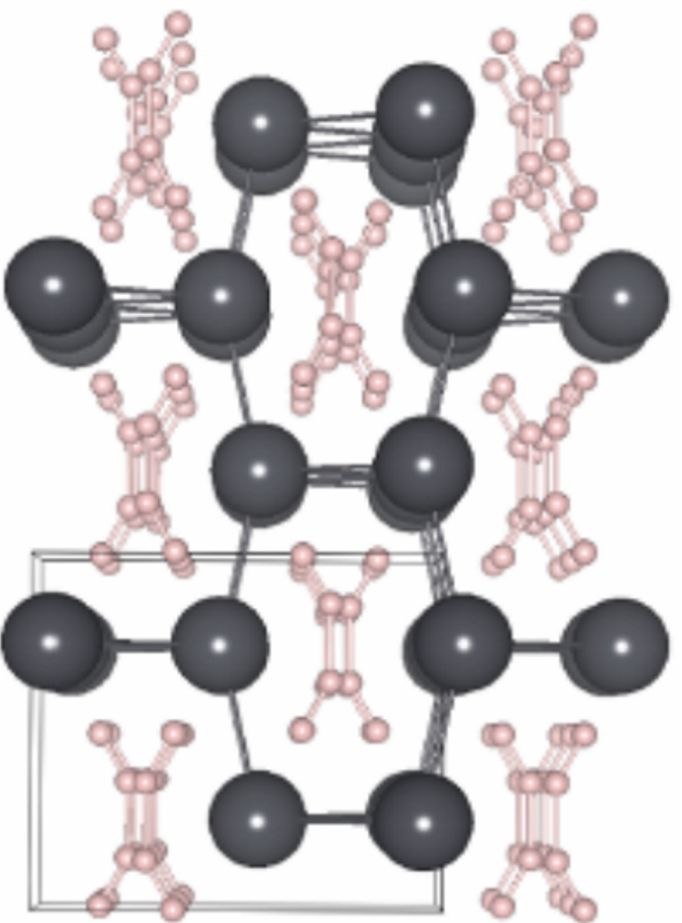
**Fig. 4.** The total density of states (DOS) calculated at 300 GPa for Phase VII (Left), and Phase VIII (Right). The dashed curves depict the densities of states for non-interacting electrons, in a three dimensional environment.

## References

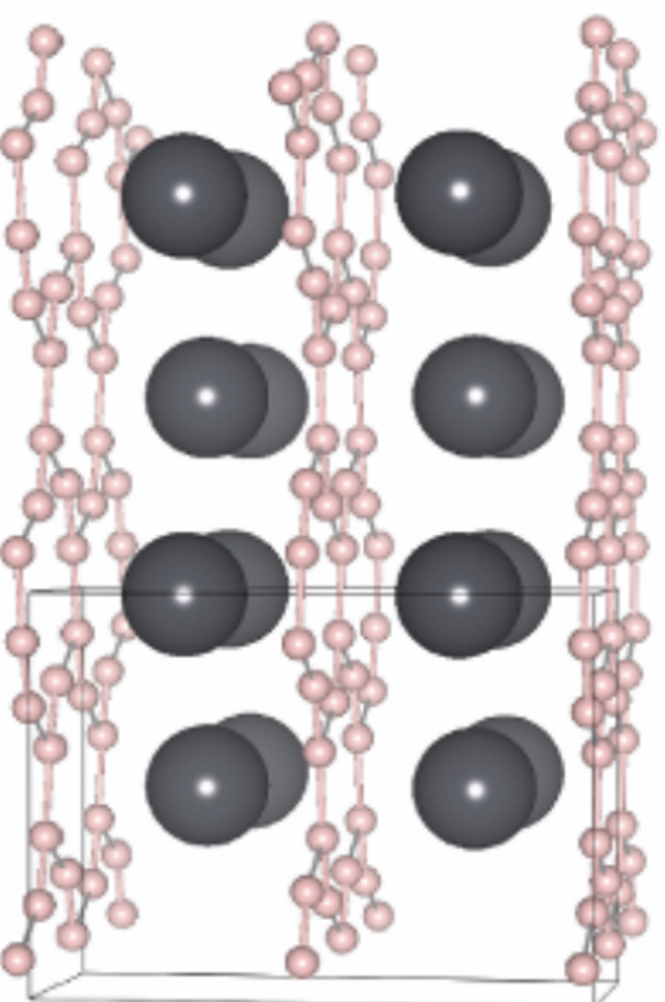
- 
1. N. W. Ashcroft, *Phys. Rev. Letters* **92**, 187002 (2004).
  2. T. Strobel, A.F. Goncharov, C.T. Seagle, Z. Liu, M. Somayazulu, V.V. Struzhkin, R.J. Hemley, *Phys. Rev. B* (in press)
  3. Z. Lia, W. Yub, C. Jina, *Solid State Comm.* **143**, 353–357 (2007).
  4. F. Paneth, E. Rabinowitsch, *Chem. Ber.* **58**, 1138 (1925).
  5. F. Paneth, *Z. Electrochem.* **26**, 452 (1920).
  6. F. Paneth, K. Fürth, *Ber. Deut. Chem. Ges.* **52**, 2020 (1919).
  7. F. Paneth, E. Rabinowitsch, *Ber. Deut. Chem. Ges.* **57**, 1877 (1924).
  8. J. P. Desclaux, P. Pyykkö, *Chem. Phys. Lett.* **29**, 534 (1974).
  9. P. Pyykkö, J. P. Desclaux, *Nature* **266**, 366 (1977).
  10. V. M. Krivtsun, Yu. A. Kuritsyn, E. P. Snegirev, *Optics and Spectroscopy* **86**, 686 (1999).
  11. X. F. Wang, L. Andrews, *J. Am. Chem. Soc.* **125**, 6581 (2003).
  12. C. W. Glass, A. R. Oganov, N. Hansen, *Phys. Comm.* **175**, 713 (2006).
  13. A. R. Oganov, C. W. Glass, *J. Chem. Phys.* **124**, 244704 (2006).
  14. D. Alfè, *Comp. Phys. Comm.* **180**, 2622 (2009).
  15. M. Martinez-Canales *et al.*, *Phys. Rev. Letters* **102**, 87005 (2009).
  16. Recomputed in our group by Dr. Vanessa Labet following the suggestions of Pickard and Needs (ref. 11).
  17. G. Gao *et al.*, *PNAS* **107**, 1317-1320 (2010).

- 
18. A. D. Becke, K. E. Edgecombe, *J. Chem. Phys.* **92**, 5397 (1990).
  19. N. E. Christensen, S. Satpathy, Z. Pawłowska, *Phys. Rev. B* **34**, 5977 (1986).
  20. G. Baym, “Lectures on Quantum Mechanics”, W. A. Benjamin Inc., (1981).
  21. P. Gonzalez-Morelos, R. Hoffmann, N. W. Ashcroft, *Chem. Phys. Chem* **11**, 3105-3112 (2010).
  22. C. Zhang *et al.*, *Europhysics Letts.* **90**, 66006 (2010).
  23. N. H. March *et al.*, *Phase Transitions* **82**, 247 (2009).
  24. G. Gao *et al.*, *Phys. Rev. Letters* **101**, 107002 (2008).
  25. D. Y. Kima *et al.*, *PNAS* **105**, 16454 (2008).
  26. M. Martinez-Canales, A. Bergara, *High Pressure Research* **26**, 369 (2006).
  27. C. J. Pickard, R. J. Needs, *Nature Phys.* **3**, 473 (2007).
  28. A. Y. Liu *et al.*, *Phys. Rev. B* **43**, 1795 (1991).



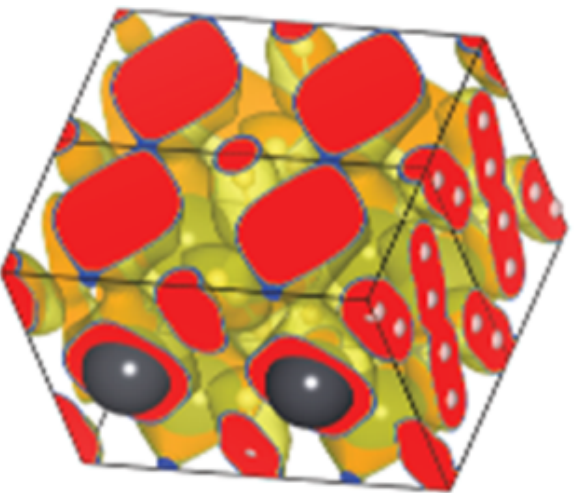


Phase VII, 250 GPa

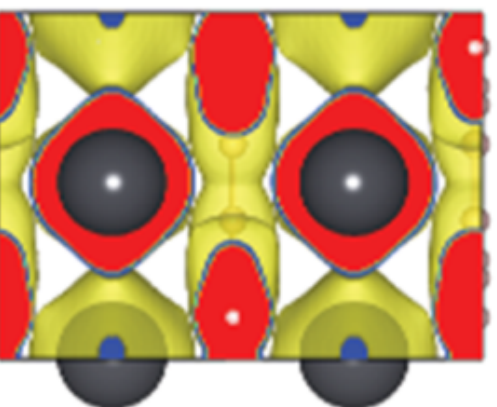


Phase VIII, 400 GPa

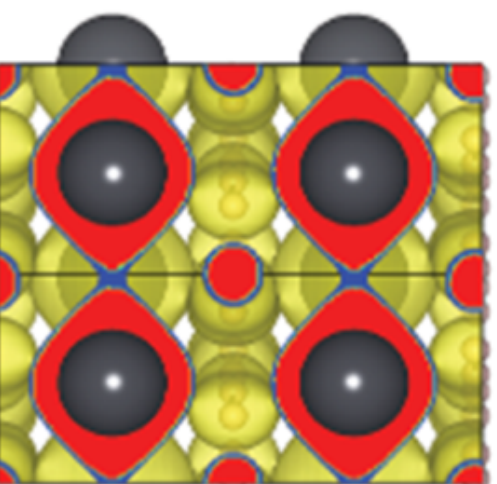




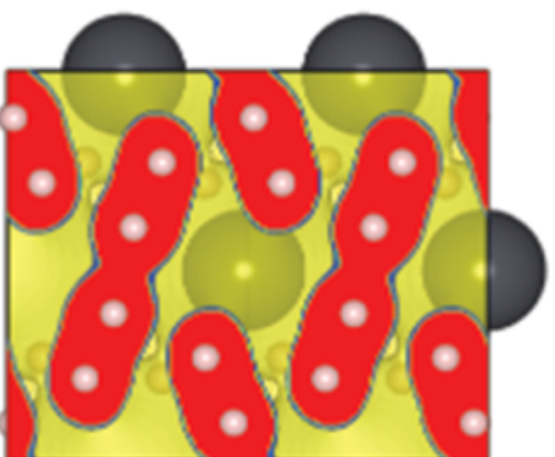
Phase VIII



Side view



Side view



Top view

