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# Computationally-Driven Experimental Discovery of the CeIr<sub>4</sub>In Compound

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We present a combined experimental and computational methodology for the discovery of new materials. Density functional theory (DFT) formation energy calculations allow us to predict the stability of various hypothetical structures. We demonstrate this approach by computationally predicting the Ce-Ir-In ternary phase diagram. We predict previously-unknown compounds CeIr<sub>4</sub>In and Ce<sub>2</sub>Ir<sub>2</sub>In to be stable. Subsequently, we successfully synthesize CeIr<sub>4</sub>In and characterize it by X-ray diffraction. Magnetization and heat capacity measurements of CeIr<sub>4</sub>In are reported. The correct prediction and discovery of CeIr<sub>4</sub>In validates this approach for discovering new materials.

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## I. INTRODUCTION

The holy grail of computational materials design is the *ab initio* discovery of new materials which can be synthesized to have specific desired properties. Our increasing understanding of the structure/properties relationship in many materials and our increasing computational resources make this an exciting possibility.<sup>1</sup> However, in many cases, the precise relationship between structure and properties is not yet well understood. This is the case for strongly correlated electron materials which exhibit interesting phenomena such as superconductivity, colossal magnetoresistance, and multiferroic behavior.<sup>2</sup>

Despite lacking a theory to immediately predict the properties of strongly correlated materials from structure alone, phenomenological observations often suggest families of composition and types of structures which are likely to exhibit specific functionalities. Having identified a family of such structures, the question becomes whether one can synthesize new materials in these families having enhanced properties. Unfortunately, the number of candidate materials which one might identify is often very large, and the experimental costs in searching for compounds can thus be heavy. Therefore, a computational method for predicting the stability of hypothetical compounds is highly desirable.

A variety of computational structure search methods have been presented including random search methods, genetic algorithms and other heuristic methods, and sophisticated data-mining approaches.<sup>3–6</sup> However, such methods are non-ideal for our application; because we are interested in searching for structures at many different compositions, random search and heuristic methods often require a prohibitively large amount of computation time. Formal data-mining approaches often fail in ternary and higher-order systems as there is an insufficient amount of data available. In addition, all of these

methods require a significant amount of resources devoted to software development. In our case, the parallel progress of experiment and theory led to fast results. Neither theory nor experiment alone could have progressed to explore this material system with modest devoted resources and on the same timescale of several months.

In this paper we report on our exploration of the Ce-Ir-In ternary phase diagram. Cerium-based transition metal indide compounds adopt a large variety of structure and bonding arrangements which produce a large variety of interesting physical phenomena including superconductivity.<sup>7</sup> Two types of structures which we are interested in exploring include systems with Ce sublattices which exist in either square planar or geometrically frustrated arrangements. Experience with cuprates, iron-pnictide, organic, and heavy fermion superconductors suggests that the square-planar compounds are likely to exhibit superconductivity.<sup>8</sup> In geometrically frustrated systems, large degeneracies can lead to novel physical phenomena.<sup>9</sup>

Using this method, we examine the structural stability of selected compounds in the Ce-Ir-In ternary phase diagram and predict two previously unknown compounds, CeIr<sub>4</sub>In and Ce<sub>2</sub>Ir<sub>2</sub>In, to be stable. Subsequently, we were able to synthesize CeIr<sub>4</sub>In, which possesses a geometrically frustrated Ce sublattice.

## II. METHODOLOGY

### A. Thermodynamics

At thermal equilibrium, a compound is stable if it has a negative formation energy with respect to its constituent elemental and any other competing phases.<sup>10</sup> The formation energy per atom for a ternary compound Ce<sub>n</sub>Ir<sub>m</sub>In<sub>l</sub> with respect to its constituent elements may be written

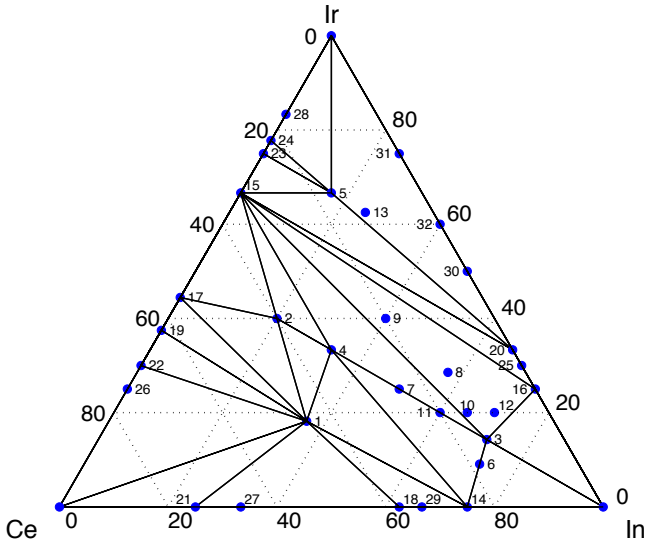


FIG. 1: (color online) Predicted Ce-Ir-In phase diagram. This top-down view of the convex hull of various calculated Ce-Ir-In phases illustrates which are stable at zero temperature. The phases at corners of each facet represent stable compounds. Additionally, some metastable phases found in this work are also plotted according to their compositions. Compounds are numbered following their order in Table I.

as

$$\Delta E_f = \frac{E_{\text{Ce}_n\text{Ir}_m\text{In}_l} - n E_{\text{Ce}} - m E_{\text{Ir}} - l E_{\text{In}}}{n + m + l}. \quad (1)$$

where  $E_{\text{Ce}}$ ,  $E_{\text{Ir}}$ , and  $E_{\text{In}}$  are the Gibbs free energies per atom of the elemental phases.

Using these formation energies, stable and competing phases can easily be illustrated. To this end, each compound is plotted according to its composition and formation energy. As shown in Fig. 1, the composition can be represented in two dimensions. Energy may be thought of as the third dimension, decreasing into the page. Then, we construct the smallest convex surface encompassing all points. This is known as the convex hull, and the thermodynamically stable structures are those which lie on the downward-facing facets of the convex hull. As the value of constructing the convex hull lies entirely in identifying these downward-facing facets, we will generally refer to them as the convex hull, neglecting the vertical and upward-facing ones.

Both our computational and experimental techniques have limitations in finding these low-energy structures. Identifying candidate structures is an open challenge to the computational methods which also have known systematic errors. The annealing process in experimental synthesis techniques may fail to find the lowest energy structure due to kinetic barriers and insufficient driving forces. Thus, synthesis routes may exist which allow us to produce metastable compounds and at the same time it may be difficult to realize the true ground state.

## B. Details of the Calculations

Calculation of the entire Gibbs free energy is computationally expensive and unnecessary. The internal energy contribution to the free energy dominates the enthalpic and entropic terms, particularly at low temperature and pressure. In this work, energies are calculated and evaluated in the zero temperature and pressure limit, neglecting zero point energies which are small in these heavy elements.

Our DFT calculations are performed using VASP<sup>11,12</sup> with the projector-augmented wave PAW method in the generalized gradient approximation PBE.<sup>13</sup> Energy differences between compounds were converged to within 0.3 meV/atom using a plane wave cutoff energy of 400 eV and a k-point mesh of density 40 Å<sup>-1</sup>. During the structural optimization, the Brillouin zone integration was performed using the Methfessel-Paxton scheme with a smearing of 0.2 eV and for the final structure, the energy was calculated using the linear tetrahedron method with Blöchl corrections.<sup>14</sup>

Our method of exploring the ternary phase space requires initial guesses of the crystal structure for a variety of compositions. The results presented here take advantage of the fact that a compound's structure is often reasonably close to that of a related ternary compound composed of different elements.<sup>15</sup> For example, the structure of CeRhIn<sub>5</sub> can be used as a guess for the structure of CeTmIn<sub>5</sub>, where TM is any transition metal. Following this approach we assembled a set of candidate structures from known compounds in the Ce-Ir-In systems and related binary and ternary systems. Specifically, we selected the known elemental, binary and ternary compounds of the Ce-Ir-In, Ce-Pt-In, Ce-Rh-In, and Ce-Pd-Al ternary and the Ru-In binary systems. We did not, however, include all known compounds of these systems, but rather selected those which experimental experience has shown most often form in Ce-Ir-In ternary phases and other chemically-similar systems.

After determining the formation energy for all the candidate structures, we determine the 3-dimensional convex hull by plotting the composition versus formation energy with respect to the elemental phases. The compounds which lie on the convex hull are energetically stable. Those above the hull have positive formation energies with respect to competing phases and are thus challenging or impossible to synthesize. From the hull, we can also identify competing phases. Calculating the formation energy of a compound with respect to competing phases allows us to quantify its degree of stability and the driving force for its formation.

The success of this project is due to the close communication between experimental and theoretical efforts. The sharing of preliminary information between the groups about particular structures and entire portions of the phase diagram was instrumental in accelerating the discovery process.

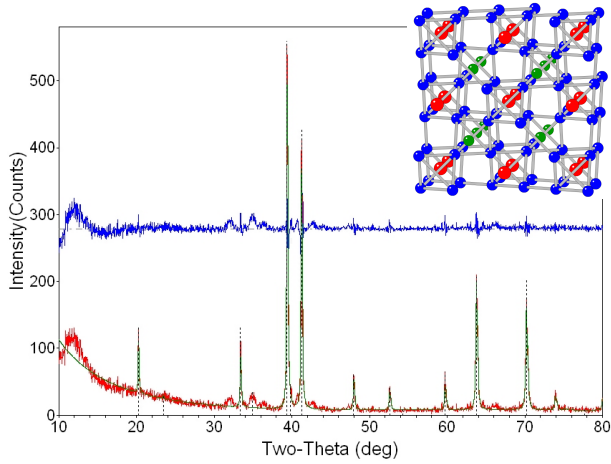


FIG. 2: (color online) X-ray powder diffraction of  $\text{CeIr}_4\text{In}$ . Tick marks indicate the expected peak positions based on the  $\text{MgCu}_4\text{Sn}$  structure type. The difference between the measured (red) and calculated (green) spectrum is shown above the measured spectrum in blue. The inset displays the crystal structure of  $\text{CeIr}_4\text{In}$  (red spheres = Ce, blue spheres = Ir, and green spheres = In).

### C. Details of the Synthesis

Compounds can be synthesized in a large variety of ways. For those compounds which were previously unknown to our knowledge (namely  $\text{CeIr}_4\text{In}$  and  $\text{Ce}_2\text{Ir}_2\text{In}$ ) we attempted to synthesize these crystals by arc melting. The synthesis of  $\text{CeIr}_4\text{In}$  was carried out via arc-melting the respective elements in a stoichiometric ratio of Ce:Ir:In - 1:4:1. The resulting button was flipped over 3 times and completely melted in order to ensure the reaction product was homogeneous. After the arc-melting, the pellet was weighed and was found to have less than a 5 mg loss. X-ray powder diffraction on the as cast pellet was completed on an X-1 Scintag diffractometer in the  $2\theta$  range of 10 - 80°. The powder pattern was refined using the JADE software package. The attempted synthesis of  $\text{Ce}_2\text{Ir}_2\text{In}$  was carried out via arc-melting the respective elements in a stoichiometric ratio of Ce:Ir:In of 2:2:1. Characterization measurements of the magnetic susceptibility and heat capacity were performed in commercial systems from Quantum Design.

## III. RESULTS

### A. Theory

For the Ce-Ir-In phase diagram we computed the formation energy of five known ternary structures ( $\text{Ce}_5\text{Ir}_2\text{In}_4$ <sup>16</sup>,  $\text{CeIrIn}_5$ <sup>17,18</sup>,  $\text{CeIrIn}_9$ <sup>19</sup>,  $\text{Ce}_2\text{IrIn}_8$ <sup>18</sup>, and  $\text{CeIrIn}_2$ <sup>19</sup>), eight currently unknown ternary structures ( $\text{Ce}_2\text{Ir}_2\text{In}$ ,  $\text{CeIr}_4\text{In}$ ,  $\text{CeIr}_2\text{In}_4$ ,  $\text{CeIr}_2\text{In}_2$ ,  $\text{CeIrIn}_3$ ,  $\text{Ce}_3\text{Ir}_4\text{In}_{13}$ ,  $\text{CeIr}_2\text{In}_7$ , and  $\text{CeIr}_5\text{In}_2$ ), nineteen binary

structures, and the three elemental phases. The results are tabulated in Table I, and the corresponding convex hull is presented in Fig. 1. Of the five known ternary compounds, three of them were computed to be stable. The two others,  $\text{Ce}_2\text{IrIn}_8$  and  $\text{CeIrIn}_2$ , had only small positive formation energies of 0.7 and 40.2 meV with respect to competing phases, respectively. Of the eight unknown ternary structures, six have positive formation energies relative to competing phases, and thus we would expect them to be difficult to synthesize. However, two of these compounds have negative formation energies with respect to competing phases. Consequently we predict two new stable compounds:  $\text{Ce}_2\text{Ir}_2\text{In}$  and  $\text{CeIr}_4\text{In}$ . Subsequent experiments described in Sec. IIIB successfully synthesized the  $\text{CeIr}_4\text{In}$ .

We characterize the electronic, magnetic and vibrational properties of the new  $\text{CeIr}_4\text{In}$  compound. Phonon calculations for the cubic  $\text{CeIr}_4\text{In}$  compound show an unstable mode and predict a rotational distortion of the Ir tetrahedra. Relaxing the crystal structure along this distortion mode rotates the tetrahedra by approximately 11° and couples to a tetragonal distortion of the lattice parameters elongating the lattice parameter parallel to the axis of rotation. The space group symmetry is lowered from cubic  $F\bar{4}3m$  (216) to tetragonal  $I\bar{4}$  (82). The lattice parameter change slightly from  $a = 7.566$  Å of the cubic structures to  $a = b = 7.508$  Å and  $c = 7.640$  Å. The symmetry breaking lowers the energy by about 300 meV per unit cell. The X-ray diffraction pattern in Fig. 2 shows no evidence of the cubic cell undergoing any such distortion. This discrepancy between calculations and experiment may potentially be due to long-range effects not captured in the limited cell size of the calculations, *e.g.*, disorder among the Ce and Ir sublattices.

Calculations of the electronic and magnetic structure of the distorted  $\text{CeIr}_4\text{In}$  compound starting from various potential magnetically ordered states always result in a non-magnetic ground states. Each Ce atom contributes one  $f$ -electron to the bonding. The  $f$ -electrons is delocalized and no magnetic moments are present. This suggests that the magnetic features seen in the experiments described in Sec. IIIB are indeed a consequence of a small amount of impurity phases.

### B. Experiment

Following the theoretical prediction of stable  $\text{Ce}_2\text{Ir}_2\text{In}$  and  $\text{CeIr}_4\text{In}$  compounds, we attempted to synthesize these compounds by arc melting. The analysis of the X-ray powder diffraction data of  $\text{Ce}_2\text{Ir}_2\text{In}$  indicates multiple competing phases which are not easily identifiable. Annealing of the  $\text{Ce}_2\text{Ir}_2\text{In}$  sample at 800°C for two weeks did not significantly alter the powder X-ray diffraction pattern. In contrast, the X-ray analysis of the arc melted  $\text{CeIr}_4\text{In}$  sample shown in Fig. 2 confirms the formation of the predicted  $\text{CeIr}_4\text{In}$  compound with lattice parameters  $a = b = c = 7.579$  Å in good agreement with the

TABLE I: Calculated energies, competing phases, and space groups associated with Ce-Ir-In phase diagram. The total energy of the compound,  $E_{\text{total}}$ , represents the calculated value per formula unit. The heat of formation of the compounds,  $\Delta E_f$ , relative to the elements and the formation energy relative to the competing phases,  $\Delta E_{\text{comp}}$ , are given units of meV/atom. Compounds listed in **bold** are known to exist.

No.	Compound [eV/f.u.]	$E_{\text{total}}$ [meV/atom]	$\Delta E_f$ [meV/atom]	$\Delta E_{\text{comp}}$	Competing Phases	Space Group
1	<b>Ce<sub>5</sub>Ir<sub>2</sub>In<sub>4</sub></b>	-64.081	-588	-65	Ce <sub>3</sub> In <sub>5</sub> , Ce <sub>5</sub> Ir <sub>3</sub> , Ce <sub>2</sub> Ir <sub>2</sub> In	<i>Pbam</i> (55)
2	Ce <sub>2</sub> Ir <sub>2</sub> In	-35.881	-749	-42	Ce <sub>5</sub> Ir <sub>4</sub> , CeIr <sub>2</sub> , CeIrIn	<i>P4/mbm</i> (127)
3	<b>CeIrIn<sub>5</sub></b>	-30.347	-395	-35	Ce <sub>2</sub> IrIn <sub>8</sub> , CeIrIn, IrIn <sub>3</sub>	<i>P4/mmm</i> (123)
4	<b>CeIrIn</b>	-19.415	-689	-32	Ce <sub>2</sub> Ir <sub>2</sub> In, CeIrIn <sub>5</sub>	<i>P62m</i> (189)
5	<b>CeIr<sub>4</sub>In</b>	-46.651	-455	-8	Ce <sub>2</sub> Ir <sub>7</sub> , IrIn <sub>2</sub>	<i>I4</i> (82)
6	<b>Ce<sub>2</sub>IrIn<sub>8</sub></b>	-45.594	-400	1	CeIn <sub>3</sub> , CeIrIn <sub>5</sub>	<i>P4/mmm</i> (123)
7	<b>CeIrIn<sub>2</sub></b>	-21.987	-520	40	CeIrIn, CeIrIn <sub>5</sub>	<i>Cmcm</i> (63)
8	CeIr <sub>2</sub> In <sub>4</sub>	-36.504	-375	44	CeIr <sub>2</sub> , IrIn <sub>3</sub> , CeIrIn <sub>5</sub>	<i>Pmna</i> (62)
9	CeIr <sub>2</sub> In <sub>2</sub>	-31.124	-472	56	CeIr <sub>2</sub> , IrIn <sub>3</sub> , CeIrIn <sub>5</sub>	<i>P21/m</i> (11)
10	Ce <sub>3</sub> Ir <sub>4</sub> In <sub>13</sub>	-93.408	-346	70	CeIr <sub>2</sub> , IrIn <sub>3</sub> , CeIrIn <sub>5</sub>	<i>Pm3m</i> (223)
11	CeIrIn <sub>3</sub>	-24.435	-394	89	CeIrIn, CeIrIn <sub>5</sub>	<i>Pm3n</i> (223)
12	CeIr <sub>2</sub> In <sub>7</sub>	-43.520	-197	140	IrIn <sub>3</sub> , CeIr <sub>2</sub> , CeIrIn <sub>5</sub>	<i>I4/mmm</i> (139)
13	CeIr <sub>5</sub> In <sub>2</sub>	-56.784	-181	189	Ce <sub>2</sub> Ir <sub>7</sub> , Ir, IrIn <sub>2</sub>	<i>P4/mmm</i> (123)
14	<b>CeIn<sub>3</sub></b>	-15.255	-412	-157	In, Ce <sub>3</sub> In <sub>5</sub>	<i>Pm3m</i> (221)
15	<b>CeIr<sub>2</sub></b>	-25.994	-782	-149	Ce <sub>5</sub> Ir <sub>4</sub> , CeIr <sub>3</sub>	<i>Fd3m</i> (227)
16	<b>IrIn<sub>3</sub></b>	-17.280	-187	-43	In, Ir <sub>3</sub> In <sub>7</sub>	<i>P4<sub>2</sub>/mnm</i> (134)
17	<b>Ce<sub>5</sub>Ir<sub>4</sub></b>	-71.542	-717	-43	Ce <sub>5</sub> Ir <sub>3</sub> , CeIr <sub>2</sub>	<i>Pnma</i> (62)
18	<b>Ce<sub>3</sub>In<sub>5</sub></b>	-33.643	-382	-30	Ce <sub>3</sub> In, CeIn <sub>3</sub>	<i>Cmcm</i> (63)
19	<b>Ce<sub>5</sub>Ir<sub>3</sub></b>	-61.358	-640	-18	Ce <sub>7</sub> Ir <sub>3</sub> , Ce <sub>5</sub> Ir <sub>4</sub>	<i>P4/ncc</i> (130)
20	<b>IrIn<sub>2</sub></b>	-14.526	-184	-18	Ir, IrIn <sub>3</sub>	<i>Fddd</i> (70)
21	<b>Ce<sub>3</sub>In</b>	-21.036	-170	-17	Ce, Ce <sub>3</sub> In <sub>5</sub>	<i>Pm3m</i> (221)
22	<b>Ce<sub>7</sub>Ir<sub>3</sub></b>	-73.297	-520	-7	Ce, Ce <sub>5</sub> Ir <sub>3</sub>	<i>P6<sub>3</sub>mc</i> (186)
23	<b>CeIr<sub>3</sub></b>	-34.913	-602	-5	Ce <sub>2</sub> Ir <sub>7</sub> , CeIr <sub>2</sub>	<i>R3m</i> (166)
24	<b>Ce<sub>2</sub>Ir<sub>7</sub></b>	-78.684	-535	0	Ir, CeIr <sub>3</sub>	<i>R3m</i> (166)
25	Ir <sub>3</sub> In <sub>7</sub>	-46.201	-172	13	IrIn <sub>2</sub> , IrIn <sub>3</sub>	<i>Im3m</i> (229)
26	<b>Ce<sub>3</sub>Ir</b>	-28.252	-399	34	Ce, Ce <sub>7</sub> Ir <sub>3</sub>	<i>Pnma</i> (62)
27	<b>Ce<sub>2</sub>In</b>	-14.929	-168	49	Ce <sub>3</sub> In, Ce <sub>3</sub> In <sub>5</sub>	<i>P63/mmc</i> (194)
28	<b>CeIr<sub>5</sub></b>	-52.303	-347	54	Ce <sub>2</sub> Ir <sub>7</sub> , Ir	<i>P6/mmm</i> (191)
29	<b>CeIn<sub>2</sub></b>	-12.050	-334	58	Ce <sub>3</sub> In <sub>5</sub> , CeIn <sub>3</sub>	<i>Imma</i> (74)
30	IrIn	-11.532	-58	80	Ir, IrIn <sub>2</sub>	<i>C2/m</i> (12)
31	Ir <sub>3</sub> In	-28.895	59	128	Ir, IrIn <sub>2</sub>	<i>Pm3m</i> (221)
32	Ir <sub>3</sub> In <sub>2</sub>	-31.479	42	152	Ir, IrIn <sub>2</sub>	<i>P31c</i> (163)
	Ce	-5.933				<i>Fm3m</i> (225)
	Ir	-8.857				<i>Fm3m</i> (225)
	In	-2.558				<i>I4/mmm</i> (139)

calculated lattice parameters.

The crystal structure of CeIr<sub>4</sub>In is presented in the inset of Fig. 2. The Ce sublattice exhibits an fcc structure composed of corner sharing tetrahedra, which satisfies the condition for geometric frustration of the magnetic moments on the Ce atoms. Consequently, we measured the magnetic susceptibility and heat capacity to determine the basic properties and whether there is evidence for magnetic frustration in CeIr<sub>4</sub>In.

The magnetic susceptibility is shown in Fig. 3. An indication for ferromagnetic ordering at low temperature is indicated by the rapid rise in the susceptibility. The inset showing  $M$  vs.  $H$  further supports ferromagnetic ordering at low temperature albeit with a small ordered moment. The inverse susceptibility shows a large region of curvature up to room temperature. Similar temperature dependence seen in SmPt<sub>4</sub>In was interpreted in

terms of a Van Vleck contribution plus crystal electric field (CEF) effects.<sup>20</sup> However, the CEF ground state for Ce in CeIr<sub>4</sub>In is most likely magnetic. Alternatively, the large region of curvature could indicate a relatively large energy scale for the onset of exchange interactions, for which the ordering between moments is suppressed to much lower temperatures.

The heat capacity data shown in the inset of Fig. 3 shows a clear albeit broad magnetic ordering transition at 2.5 K consistent with the susceptibility data and resistivity measurements (not shown) and a smaller anomaly at 6 K. The field dependence of the anomaly at 2.5 K (not shown) and the shape of the transition in susceptibility again suggests that the magnetic ordering is ferromagnetic. The extrapolated zero temperature Sommerfeld coefficient in the ordered state is 55 mJ/(mol·K<sup>2</sup>). From the specific heat data, we calculate the entropy

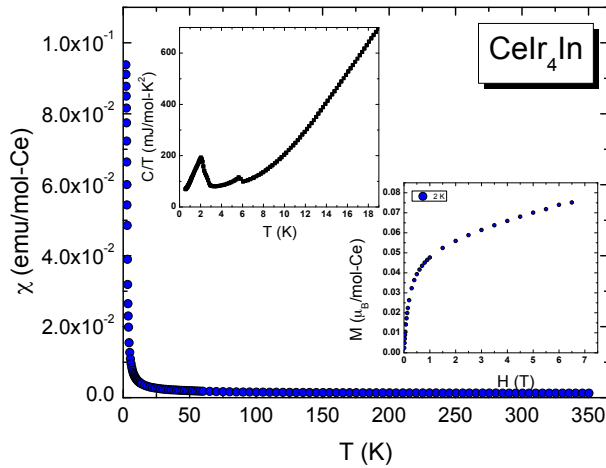


FIG. 3: (color online) Magnetic susceptibility of  $\text{CeIr}_4\text{In}$  versus temperature. Upper left inset displays the specific heat of  $\text{CeIr}_4\text{In}$  versus temperature and the lower right inset shows the magnetization as a function of magnetic field at 2 K.

$S = \int C/T dT$ . However, only about 5% of  $R \ln 2$  in entropy is recovered at the transition temperature, much smaller than expected for a system of local magnetic moments at the ferromagnetic ordering temperature.

The small amount of entropy and small magnetic moment inferred from the susceptibility measurements suggests that these features are a consequence of a small amount of impurity phases. We note that there are small unaccounted peaks in the X-ray pattern shown in Fig. 2. While we believe this is the most likely explanation a few alternative scenarios still exist. First, the reduced entropy may be a consequence of Ce moments having largely itinerant character (*i.e.* in a  $\text{Ce}^{4+}$  configuration) as has been suggested to occur in  $\text{CeNi}_4\text{In}$ .<sup>21</sup> Alternatively, a large majority of the magnetic entropy may be quenched at much higher temperatures as a result of disorder-, RKKY-, and/or geometric-induced magnetic frustration effects as has been suggested to occur in the  $\text{RCu}_4\text{In}$  ( $R = \text{Gd-Er}$ ) compounds.<sup>22</sup> Finally, Kondo screening is another alternative to quenching the magnetic entropy at higher temperatures without long range order.

Of all the  $\text{CeTM}_4\text{In}$  compounds,  $\text{CePt}_4\text{In}$  is most similar in structure and chemistry to  $\text{CeIr}_4\text{In}$ .  $\text{CePt}_4\text{In}$  is a local moment system displaying heavy fermion behavior<sup>23,24</sup>, with a Kondo temperature of  $\sim 10$  K and short range ordering occurring at 250 mK<sup>25</sup> which would favor some combination of the latter two scenarios for explaining the reduced entropy in  $\text{CeIr}_4\text{In}$ . If Ir hybridizes more strongly than Pt, this would favor the scenario where the magnetic anomalies are a result of impurity phases, and  $\text{CeIr}_4\text{In}$  is approaching mixed valent behavior akin to  $\text{CeNi}_4\text{In}$ . Consequently, it is currently inconclusive as to whether geometric frustration is playing a significant role in  $\text{CeIr}_4\text{In}$ . More measurements are needed to determine the origin of the physical properties in  $\text{CeIr}_4\text{In}$ .

#### IV. DISCUSSION

In retrospect, it may not be surprising to anticipate the discovery of  $\text{CeIr}_4\text{In}$  when the isostructural analog  $\text{CePt}_4\text{In}$  is known. However, all of the unknown Ce-Ir-In compounds computed in Table I contain closely related known isostructural analogs. Furthermore, it is remarkable that these relatively simple DFT calculations which neglect zero point motion, finite temperature, and electronic correlation effects, accurately predict the existence and lack thereof, for 10 out of the 13 ternary compounds and 15 of the 19 binary compounds investigated here in the Ce-Ir-In ternary phase diagram. In addition, for  $\text{Ce}_2\text{IrIn}_8$  and  $\text{CeIrIn}_2$  which are predicted to be unstable but are known, and  $\text{Ce}_2\text{Ir}_2\text{In}$  which is predicted to be stable but has yet to be found, the formation energies relative to competing phases are sufficiently small that we might expect that some of the neglected effects mentioned above might change the conclusion with regards to compound formation. We must then also acknowledge that given the small positive formation energy with respect to competing phases for  $\text{CeIr}_2\text{In}_4$  and  $\text{CeIr}_2\text{In}_2$ , which is less than 56 meV/atom, it would not be surprising if either or both of these compounds are eventually synthesized.

In addition to the possible shortcomings of the DFT calculations mentioned earlier, we also recognize that the phase space of potential compounds which we are exploring is remarkably small. Not only have we selected a relatively small number of compositions, but we have also assumed a particular structure type for each of these compositions. In this respect, the addition of simulated annealing, genetic algorithms and other approaches<sup>26</sup> to this approach may aide in discovering additional structures not considered based on analogy.

The fact that arc melting  $\text{Ce}_2\text{Ir}_2\text{In}$  and annealing did not produce the anticipated structure does not necessarily mean that the compound is unstable. However, it shows that  $\text{Ce}_2\text{Ir}_2\text{In}$  can not be a congruently melting compound and that the phase is either not stable at 800°C or that a direct transformation from the competing phases may exhibit slow kinetics. Alternative synthesis routes should be considered to stabilize the compound. These could be as simple as using a different temperature annealing profile or using another synthesis route such as flux growth.<sup>27</sup> Thus, we anticipate that the effort in trying to synthesize a theoretically predicted compound by this method will be commensurate with the interest that the expected properties of that compound would generate.

Despite these potential complications, the fact that our combined computational and experimental approach demonstrates an accuracy of better than 80% in identifying which compounds exist or not, proves that this can be a useful method for computationally identifying new stable structures.

## V. SUMMARY

We present a simple method combining density-functional calculations with arc-melting synthesis to efficiently investigate the structural stability of hypothetical compounds with the aim to improve the search for new compounds with desired structures and properties. As a proof of principle we applied our approach to the Ce-Ir-In phase diagram and predicted two previously unknown compounds,  $\text{Ce}_2\text{Ir}_2\text{In}$  and  $\text{CeIr}_4\text{In}$ , to be stable. Subsequent attempts to synthesize these compounds led to the discovery of  $\text{CeIr}_4\text{In}$ . We expect that this approach will lead to the discovery of many more interesting compounds with specifically desired properties.

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- <sup>1</sup> P. Piquini, P. A. Graf and A. Zunger, Phys. Rev. Lett. **100**, 186403 (2008).
  - <sup>2</sup> V.I. Anisimov, *et al.* J. Phys. Condens. Matter **9**, 7359 (1997); E. Dagatto, Science **309**, 257 (2005);
  - <sup>3</sup> *Modern Methods of Crystal Structure Prediction*, edited by A. R. Oganov, Wiley-VCH, Weinheim, Germany (2010).
  - <sup>4</sup> O. Levy, Gus L. W. Hart and S. Curtarolo, J. Am. Chem. Soc. **132**, 4830 (2010).
  - <sup>5</sup> C. Fischer, K. Tibbetts, D. Morgan, G. Ceder, Nature Materials, **5**, 641 (2006).
  - <sup>6</sup> A. van de Walle, M. Asta and G. Ceder, CALPHAD Journal **26**, 539, (2002).
  - <sup>7</sup> Ya. M. Kalychak, *et al.*, Handbook on the Physics and Chemistry of Rare Earths, **34**, 1 (Chapter 218) (2005).
  - <sup>8</sup> P. Monthoux, D. Pines, and G.G. Lonzarich, Nature **450**, 1177 (2007).
  - <sup>9</sup> A.P. Ramirez, Annu. Rev. Mater. Sci. **24**, 453 (1994).
  - <sup>10</sup> Y. A. Chang and W. A. Oates, *Materials Thermodynamics*, John Wiley & Sons, Hoboken (2010).
  - <sup>11</sup> G. Kresse and J. Furthmüller, Phys. Rev. B **54**, 11169 (1996).
  - <sup>12</sup> G. Kresse and D. Joubert, Phys. Rev. B **59**, 1758 (1999).
  - <sup>13</sup> P. E. Blöchl, Phys. Rev. B **50**, 17953 (1994).
  - <sup>14</sup> P. E. Blöchl, O. Jepsen, and O. K. Andersen, Phys. Rev. B **49**, 16223 (1994).
  - <sup>15</sup> R. G. Hennig, A. E. Carlsson, K. F. Kelton, and C. L. Henley, Phys. Rev. B **71**, 144103 (2005).
  - <sup>16</sup> R. Zaremba, W. Hermes, M. Eul, and R. Pöttgen, Z. Naturforsch. **63b**, 1447 (2008).
  - <sup>17</sup> C. Petrovic *et al.*, Europhys. Lett. **53**, 354 (2001); E.G. Moshopoulou, Z. Fisk, J.L. Sarrao, and J.D. Thompson, J. Solid State Chem. **158**, 25 (2001)
  - <sup>18</sup> J.D. Thompson *et al.*, J. Mag. Mag. Mat. **226**, 5 (2001).
  - <sup>19</sup> V.I. Zaremba, U.Ch. Rodewald, and R. Pöttgen, Z. Anorg. Allg. Chem. **63**, 1065 (2005).
  - <sup>20</sup> S.K. Malik, *et al.*, J. Mag. Mag. Mat. **92**, 80 (1990).
  - <sup>21</sup> N. Tsujii, *et al.*, Physica B, **329**, 651 (2003).
  - <sup>22</sup> V. Fritsch, J. D. Thompson, and J. L. Sarrao, Phys. Rev. B **71**, 132401 (2005).
  - <sup>23</sup> S. K. Malik, D. T. Adroja, M. Slaski, B. D. Dunlap, and A. Umezawa, Phys. Rev. B **40**, 9378 (1989).
  - <sup>24</sup> A. P. Pikul *et al.*, Phys. Rev. B **73**, 092406 (2006).
  - <sup>25</sup> A. P. Pikul *et al.*, Physica B **403**, 842 (2008).
  - <sup>26</sup> S. M. Woodley and R. Catlow, Nature Mat. **7**, 937 (2008).
  - <sup>27</sup> Z. Fisk and J.P. Remeika, Handbook on the Physics and Chemistry of the Rare Earths, ch. 81, vol. 12, eds. K.A. Gschneidner and L. Eyring, p. 53 (1989).